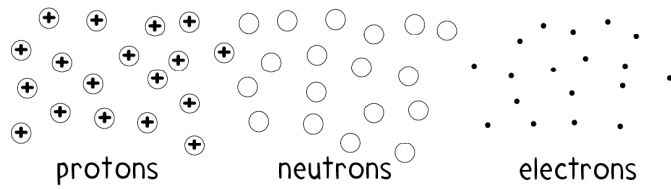


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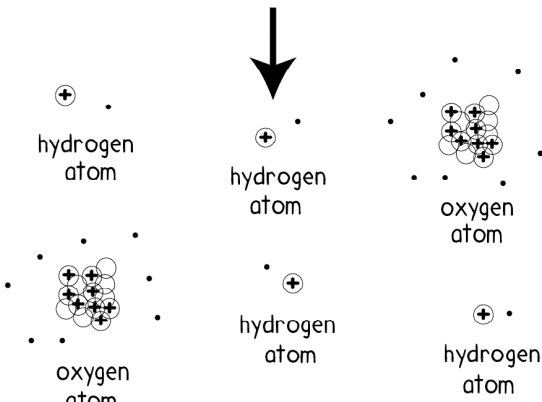
Chapter 2: The Chemistry of Life

Atoms to Molecules to Molecular Attractions



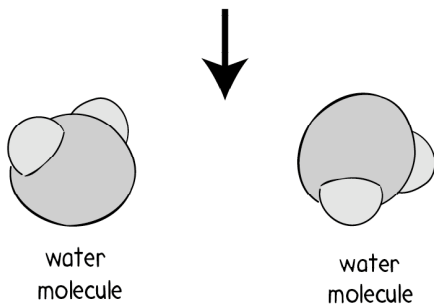
Subatomic particles are the fundamental building blocks of all _____.

SUBATOMIC PARTICLES



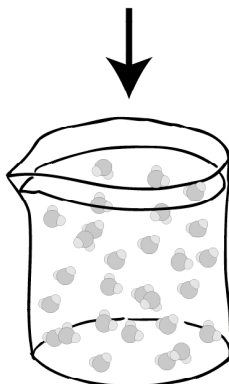
An atom is a group of _____ held tightly together. An oxygen atom is a group of 8 _____, 8 _____, and 8 _____. A hydrogen atom is a group of only 1 _____ and 1 _____.

ATOMS



A _____ is a group of atoms held tightly together. A water _____ consists of 2 _____ atoms and 1 _____ atom.

MOLECULES



WATER

Water is a material made up of billions upon billions of water _____. Water's physical properties are based upon how these water _____ interact with one another.

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Chapter 2: The Chemistry of Life Solutions

1. Use these terms to complete the following sentences. Some terms may be used more than once.



- | | | |
|-----------|---------------|----------------|
| solution | solvent | solute |
| dissolve | concentrated | dilute |
| saturated | concentration | mole |
| molarity | solubility | soluble |
| insoluble | precipitate | supersaturated |

Sugar is _____ in water, for the two can be mixed homogeneously to form a _____. The _____ of sugar in water is so great that _____ homogeneous mixtures are easily prepared. Sugar, however, is not infinitely _____ in water, for when too much of this _____ is added to water, which behaves as the _____, the solution becomes _____. At this point any additional sugar is _____ for it will not _____. If the temperature of a saturated sugar solution is lowered, the _____ of the sugar in water is also lowered. If some of the sugar comes out of solution, it is said to form a _____. If, however, the sugar remains in solution despite the decrease in solubility, then the solution is said to be _____. Adding only a small amount of sugar to water results in a _____ solution. The _____ of this solution or any solution can be measured in terms of _____, which tells us the number of solute molecules per liter of solution. If there are 6.022×10^{23} molecules in 1 liter of solution, then the _____ of the solution is 1 _____ per liter.